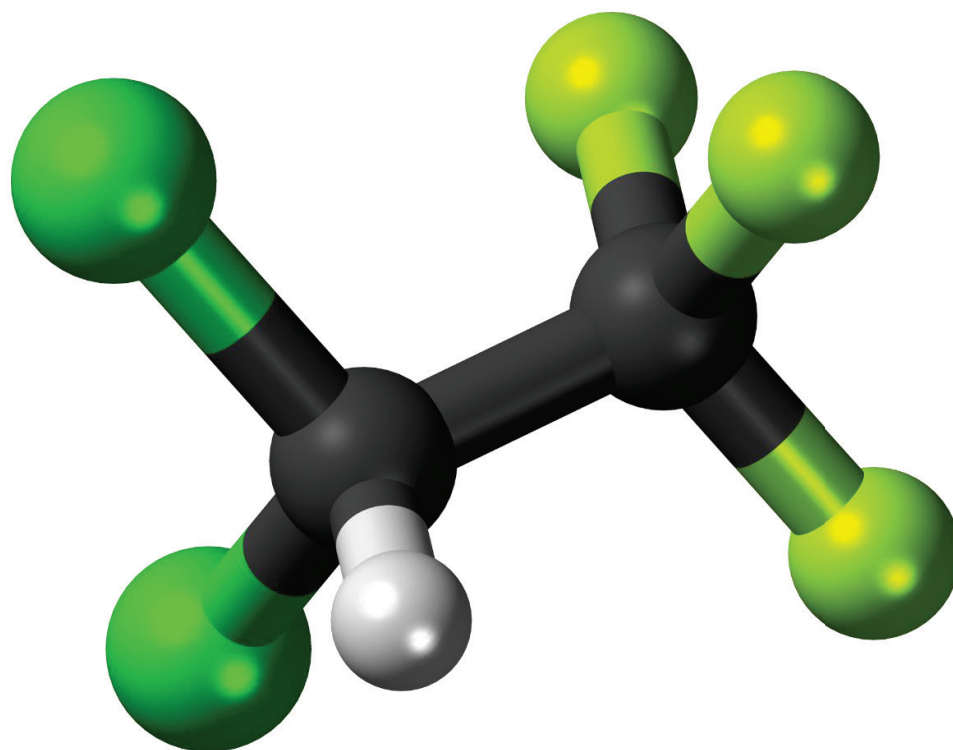


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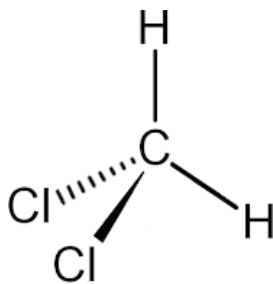
Polyhalogen Compounds

Polyhalogen compounds: Carbon compounds containing more than one halogen atom permolecule.

Polyhalogen compounds are useful in various industries and in agriculture.
Some important polyhalogen compounds:



Dichloromethane (Methylene chloride)



Uses:

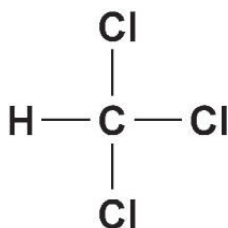
Dichloromethane (methylene chloride) is used as a:

1. Solvent for paint removers
2. Propellant in aerosols
3. Process solvent in the manufacture of drugs
4. Metal cleaning and finishing solvent

Harmful effects:

1. It endangers the human central nervous system.
2. Exposure to lower levels of methylene chloride in air can lead to slightly impaired hearing and vision.
3. High levels of methylene chloride in air cause dizziness, nausea, tingling and numbness in the fingers and toes.
4. In humans, direct skin contact with methylene chloride causes intense burning and mild redness of the skin.
5. Direct contact with the eyes can burn the cornea.

Trichloromethane (Chloroform)



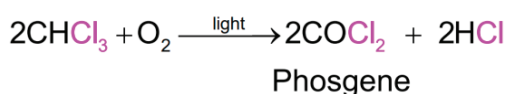
Uses:

1. Chemically, chloroform is used as a solvent for fats, alkaloids, iodine and other substances.
2. The major use of chloroform today is in the production of the freon refrigerant R-22.

3. It was once used as a general anaesthetic in surgery but has been replaced by less toxic, safer anaesthetics such as ether.

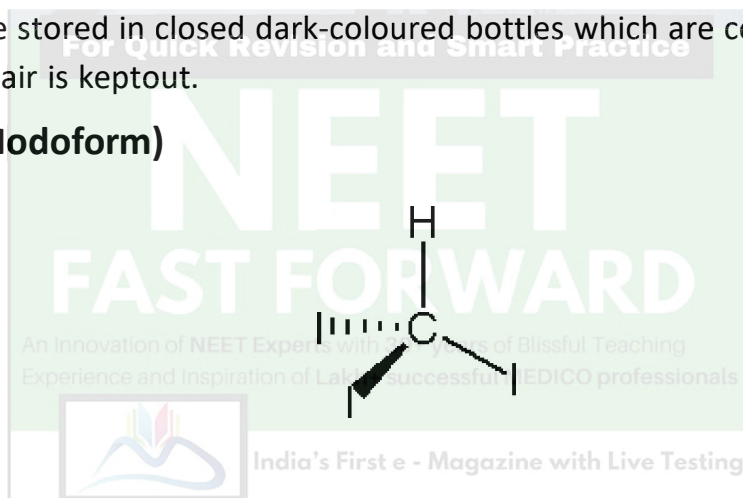
Harmful effects:

1. As might be expected from its use as an anaesthetic, inhaling chloroform vapour depresses the central nervous system.
2. Breathing about 900 parts of chloroform per million parts of air (900 ppm) for a short time can cause dizziness, fatigue and headache.
3. Chronic chloroform exposure may cause damage to the liver (where chloroform is metabolised to phosgene) and to the kidneys. Some people develop sores when the skin is immersed in chloroform.
4. Chloroform is slowly oxidised by air (oxygen) in the presence of light to an extremely poisonous gas, carbonyl chloride, also known as phosgene.



It is therefore stored in closed dark-coloured bottles which are completely filled so that air is kept out.

Triiodomethane (Iodoform)



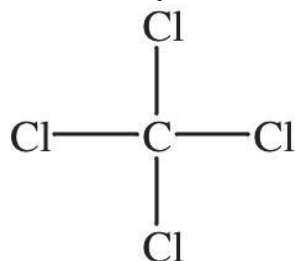
Uses:

- It was used earlier as an antiseptic, but the antiseptic properties are due to the liberation of free iodine and not due to iodoform itself.

Drawback:

- Because of its objectionable smell, it has been replaced by other formulations containing iodine.

Tetrachloromethane (Carbon tetrachloride)





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